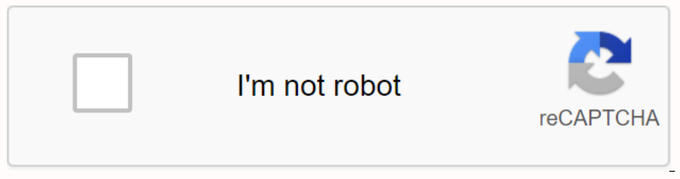

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Alcohols phenols and ethers class 12 notes pdf download

Ethers are classified by the oxygen atom-related group. Phenols and alcohols are classified by two factors: the number of hydroxyl groups based on the hybridisation of carbon atom, to which the -OH group CBSE class 12 Chemistry, 11 Alcohols, phenols and ethers – related relationships Alcohols can be made by the hydration process of alkenes or by carbonyl compounds. Hydration of alkeens – When preparing this method is either the presence of acid or hydroboration-oxidation reaction. Carbonyl compounds – This method is manufactured by catalytic reduction, as well as by the action of Grignard reagents. Preparation of phenols Alcohols can be prepared by replacement reaction or hydrolysis of diazonium salts or caraway replacement reaction – It can be prepared by two processes, namely sulphonic acid in aryylsulfonic acids, by the -OH group and halogenated atoms. Ethers can be prepared for either dehydration of alcohols and Williamson synthesis. Hydrogen halides can leave c-o bandages in the ethers. Both alcohols and phenols are inherently acidic. For more information about dehydration alcohols, see below Video: Some important questions about IUPAC names of compounds are presented below. Write their structure Cyclopent-3-en-1-ol 2-ethoxy-3-methylpentan 2-methylpentanne 2-methylbutet-2-ol 4-chloro-3-ethylbutane-1-ol Propanol boiling point is greater than the boiling point of carbon dioxide, butane. Explain. Cyclohexayl methanol synthesis by alkyl halide in the SN2 reaction. Explain Kolbe by example of a reaction. What is more acidic between ortho nitrophenol and orthomethoxyphenol? Why? Ethanol has a higher boiling point than methoxymethane? Explain. What is the equation of reactions below? Nitration anisole Friedel-Craft's acetylation anisole To learn more about this chapter and alcohol phenols and ethers class 12 Important issues to register byju's! Other important links: phenole electrophilic replacement Ether Preparation Class 12 Chemistry Notes Group 11 Alcohols Phenols Ethers Class 12 Chemistry Notes Group 11 Alcohols Phenols Ethers. PDF download for free. The tutorial's complete guide to CBSE StudentsNCERT Solutions, NCERT Exemplars, Revison Notes, Free Videos, CBSE Papers, MCQ Tests and more. Download NowAlcohols Phenols and Ethers Class 12 Notes Chemistry PDF is available for free to download myCBSEguide mobile app. The best app for CBSE students now offers alcohol phenols and ethers in Class 12 Notes the latest chapter in smart notes for the quick preparation of CBSE board exams and school-based annual exams. Class 12 Chemical Notes chapter 11 Alcohols Phenols and ethers are also available for download on the CBSE Guide website. Alcohols Phenols and Ethers Class 12 Notes Chemical Download CBSE Class 12th revision notes Group 11 Alcohols Phenols and Ethers in PDF format free of charge. 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Alcohol boiling points decrease with an increase in the branching of the carbon chain. This is due to the reduction of the forces of van der Waals with the reduction in surface area.b) Solubility: the solubility of alcohols and phenols dissolves in water because they have the ability to form hydrogen bonds with water molecules. Alcohol solubility decreases with an increase in the size of alkyl/ayl groups (hydrophobic) groups.I. O-H bond cleavage reactions: alcohols react as nucleophilia:a) Reaction with metals (b) Esterisation reaction II. Alcohol reactions accompanied by carbon cleavage – oxygen (C-O) bond: a) Reaction with hydrogen halides) Reaction of phosphorus trihalidesc) Reaction to dehydration reaction). Oxidation reaction(ii) (ii) Chemical properties of phenols:I. O-H bandage cleavage reactions: alcohols react as nuclofiphile:a) Reaction with metals (b) estering reactionII. Other chemical reactions of phenols:III. Acid nature of phenol and alcohol:a) Phenol > H2O > Primary alcohol > Secondary alcohol > tertiary alcohol. The alcoholic acidity is due to the polarity of the O-H bond. The alkyl group is an electron-releasing group (-CH3, C2H5) or has an electron-releasing inductive action (+I effect). Due to the effects of alkyl groups, the density of electrons increases to oxygen. This reduces the polarity of the O-H bond. And that's where the acid strength is decreasing. (b) Phenol is more acidic than alcohol: in phenol, the hydroxyl group is directly attached to the carbon of benzene ring sp2-hydrobriifized, which acts as an electron extraction group, while in the case of alcohols the hydroxyl group is attached to an alkyl group with a reductive effect on electrons. In phenol, the hydroxyl group is directly attached to the carbon hybridised with benzene ring sp2, while in alcohol the hydroxyl group is attached to the alkyl group sp3hybridized carbon. The electron egatiness of sp2 hybrid carbon is greater than Carbon. Thus, the polarity of the O-H bond of phenols is higher than that of alcohol. Thus, ionization of phenols is higher than that of alcohol. Alcohol and phenol ionisation is as follows: In the alcocoxibion, the negative charge is localized for oxygen when there is a charge in the phenoxideone. The delocation of the negative charge makes the phenoxide ion more stable and promotes ionization of phenol. Although phenol also has a repositioning of charges, its resonance structures have a loan ing rate, which is why the phenol molecule is less stable than phenoxide ions. (c) In substituted phenols, the presence of electron extraction groups, such as the nitro group, increases the acid strength of phenol. On the other hand, electron-releasing groups, such as alkyl groups, in general, reduce acid strength. This is because electron extraction groups lead to effective relocation of the negative charge in phenoxide ions.(a) Alcohols and phenols Phenol when responding to neutral FeCl3 gives a purple colour, while alcohols do not give a purple colour.b) Lucas reagents for primary, secondary and tertiary alcohols: if alcohol is the primary alcohol primary, there is no turbidity at room temperature. Turbidity occurs only during heating. If secondary alcohol is involved, cloudiness occurs after 5 minutes. In the case of tertiary alcohol, a test for Methanol and ethanol iodoform occurs immediately: ethanol when reacted (I2 and NaOH) or NaOI gives yellow iodoform because it is present in the CH3-CH (OH) group. Structure of the eths: Manufacture of ethers:(a) Alcohols from alkyl halide and sodium alchoxide Here, alkylhalide must be primary and alcocoxib must be on tertiary gas. In the case of aromatic ether, the aromatic part should be combined with phenoxide ion.a) Miscibility: the mixture of ethers with water is similar to that of alcohols of the same molecular weight. This is due to the fact that, like alcohols, oxygen ether can also form hydrogen bonds with a water molecule.b) Boiling points: Ethers have a much lower boiling point than alcohols. This is due to the binding of hydrogen in alcohols. Hydrogen binding ethers.a) C-O bandage cleavage in ethers: R-O-R' + HX → R-X + R'OH Excess hydrogen halides reactivity is as follows: HI >HBr>Alkyl halogens formed by HCl are always lowered with subkyl halides. But when the tertiary alkyl group is present, alkylhalogen is always tertiary. In the case of phenol ethers, cleavage occurs with phenol and alkylhalod.b) Electrophilic replacement reaction in aromatic ethers:The electrophilic replacement reaction of the aromatic ether includes the following reaction:Other conversion reactions:a) phenol to salicyl dehyde) Phenol benzene diazonium chlorideCBSE class-126s and main points Phenols and Ethers Class 12 Notes. CBSE's quick review note on class-12 chemistry, physics, mathematics, biology and other topic is very useful to watch across syllabus exam days. The correction notes shall cover all the essential examples and definitions set out in the chapter. Even if you want an overview of the chapter, quick revision notes are here for what to do if for you. These notes will definitely save you time during stressful exam days. 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